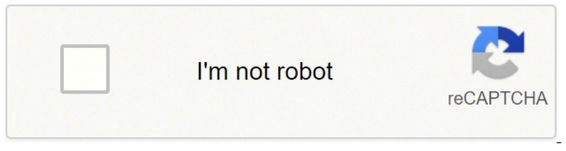

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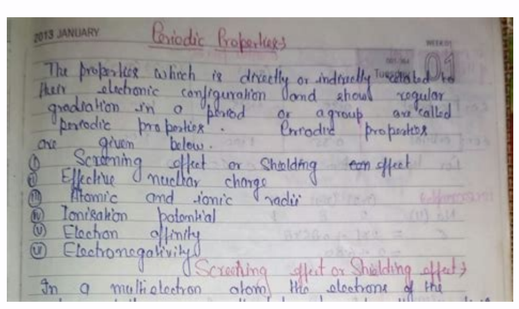
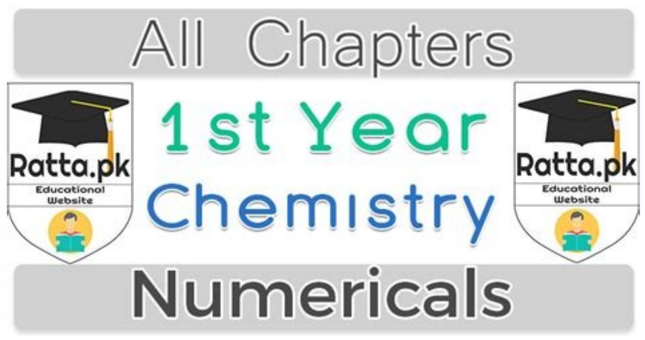
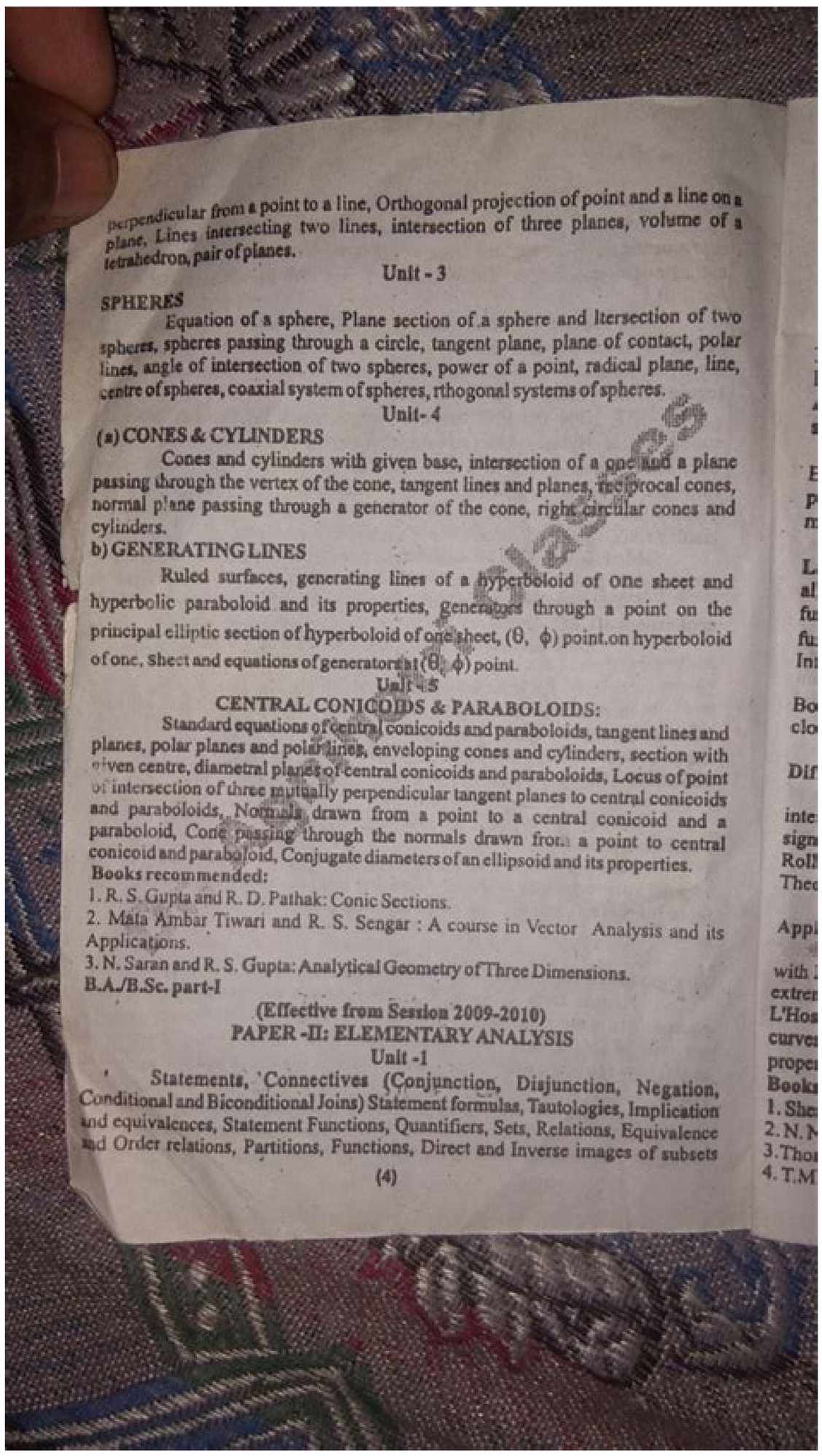
Bsc 1 year chemistry notes pdf






perpendicular from a point to a line, Orthogonal projection of point and a line on a plane, Lines intersecting two lines, intersection of three planes, volume of a tetrahedron, pair of planes.

Unit - 3

SPHERES

Equation of a sphere, Plane section of a sphere and Itersection of two spheres, spheres passing through a circle, tangent plane, plane of contact, polar lines, angle of intersection of two spheres, power of a point, radical plane, line, centre of spheres, coaxial system of spheres, rthogonal systems of spheres.

Unit- 4

(a) CONES & CYLINDERS

Cones and cylinders with given base, intersection of a one and a plane passing through the vertex of the cone, tangent lines and planes, reciprocal cones, normal plane passing through a generator of the cone, right circular cones and cylinders.

b) GENERATING LINES

Ruled surfaces, generating lines of a hyperboloid of one sheet and hyperbolic paraboloid and its properties, generators through a point on the principal elliptic section of hyperboloid of one sheet, (θ, ϕ) point on hyperboloid of one, sheet and equations of generators at (θ, ϕ) point.

Unit -5

CENTRAL CONICOIDS & PARABOLOIDS:

Standard equations of central conicoids and paraboloids, tangent lines and planes, polar planes and polar lines, enveloping cones and cylinders, section with given centre, diametral planes of central conicoids and paraboloids, Locus of point of intersection of three mutually perpendicular tangent planes to central conicoids and paraboloids, Normals drawn from a point to a central conicoid and a paraboloid, Cone passing through the normals drawn from a point to central conicoid and paraboloid, Conjugate diameters of an ellipsoid and its properties.

Books recommended:

1. R. S. Gupta and R. D. Pathak: Conic Sections.
2. Mata Ambar Tiwari and R. S. Sengar : A course in Vector Analysis and its Applications.
3. N. Saran and R. S. Gupta: Analytical Geometry of Three Dimensions.

B.A./B.Sc. part-I

(Effective from Session 2009-2010)

PAPER -II: ELEMENTARY ANALYSIS

Unit -1

Statements, Connectives (Conjunction, Disjunction, Negation, Conditional and Biconditional Joins) Statement formulas, Tautologies, Implication and equivalences, Statement Functions, Quantifiers, Sets, Relations, Equivalence and Order relations, Partitions, Functions, Direct and Inverse images of subsets

(4)

Inert pair effect causes variable valence of elements. As you will explain any three of the following items: i. Body centered 6. The modified form of the apparatus is shown below consists of a sealed fuel chamber, called a pump, containing a heavy amount of the substance in a dish, along with oxygen below about 20 atm. It is also surrounded by an

outer coating to ensure complete isolation from the atmosphere. R = (M/D) X ÃÃ˜·1/8 If ÃÃ˜Â· = 1 then M/D = R So we can say that Rheochor is the molar volume of the liquid at the temperature at which its unit viscosity. The waste is extracted with water, boiled with a little nitric acid and then a hot solution of ammonium molybdate is added to it in
excess. The energy of the ninth level of energy is given as – E = nhÃ˜ 1ã˜˜˜2 where ã˜n' is integer with values 0,1,2,3... Each electron moving into a molecular orbital has a spin of +1/2 and -1/2. B2H6 + 6H2O = 2H3BO3 + 6H24. Then the nucleus pulls the “Electron” strongly. Oxides of the phosphorus ii. What is the unit of dipole moment 4. Now,
knowing the weight of the organic compound and BaSO4, the percentage of sulfur in the organic compound is easily calculated. A. d. at ordinary temperature, d is very small in comparison with D then M to 31 /4/D = P If Ã3 = 1 at a particular temperature, then M/D = P Thus, at a particular temperature, the molar volume of a liquid having a unit of
surface tension is called Parachor. The various factors on which the enthalpy of the reaction depends are given below: The enthalpy of the reaction changes with the change in the physical state, because as the physical state changes, the extent of heat evolves. It is prepared by the action of H3PO4 on Mg-boride-H3PO4 H3PO4 + Mg3B2 = B2H6 +
Mg3 (PO4) 2 2. Acidic acid Molar volume of any substance is the volume occupied by a mole of the substance. It is a fast process. The law of rational indices states that the µ intercepts of any face of a crystal along the crystallographic axes are equal to the µ intercept or some multiples in a simple integer of them. The temperature at which the Joule
Thomson coefficient becomes zero is called the Inversion Temperature (Ti). We know that the Einstein equation - E = mc2 — (eq.1) where, E = energy, m = mass of the particles and c = the speed of light According to the day theory of Plank- E = hÃ1â µ 2 = h.c/Ã 2 = h.c/Ãτâ³‡, or, mc = h/Ãτ³‡λor, p = h/ÃTTTTTTTTTO [as mass (m)
velocity X (c) = momentum (p)] or, Ãτ‡TTTO = h/p — (eq.3) Equation âTOTO3'. Distinction between 1st, 2nd and 3rd Alcohols. HF and HCl iii. is a powerful oxidizing agent. C60 is not over µ Inc. 5. O2+, O2, and O2- Identify the sppA ©cies that are paramagnH between the above, in ascending order of alloy length. Explain any two of the following:
i. The ionization potential of Na is less than Mg iii. Each molecular orbital has been described by a wave function (Ì) called molecular orbital wave function. When a system changes from one state to another, the operation is called a process. Define the ionization potential and explain why? The Bond Dissociation energy can be defined as, The amount
of energy needed to break a molecular bond present between the Utoms of a gaseous mol. the volume so expressed in ml. In 1924, Sugden modified the above equation as - H31/4/ (D Ã¢ Â ÂÂ Âd) = MC = P where M = the molecular weight of the liquid and P = the parachor. a. These points are known as a trellishing point or trellishing site. Hess's
law finds its application seµÃ§Ãaer seµÃ§Ãaer arap saŞÃnadum sad rolac od oÉÃ§Ãanimreted fo erutcurts eht ebircseD .etihparg dna dnomaid fo ecnereffid erutcurts eht eviG .krow fo tnuoma tnelavique na otni ygrene taeh eht trevnoc ot elbissop ton si ti yllacitcarP .ylgnorts tca snortcele-'s' tsom retudo no egrahc raelcun evitceffe eht ,slatibro 'f' ro 'd'
fo tceffe gnidleihs roop ot eud dna snortnordna snortnortdna snortdna cele 'f' ro 'd' renni niatnoc stnemele eseht taht ees ew ,tnemegnarra cinortcele rieht morF .silakla sa llew sa sdica ni elbulos si OnZ tub ylno sdica ni elbulos si OaC 2H + 2X + 2)HO(M >— O2H2 + 2XM — 2H + -HO2 >— e2 + O2H2 e2 + 2X >— -X2 -X2 + 2+M >— 2XM .ssecorp
lamrehtoreht si dellac si erutarepmet tnatsnoc ta ecalp sekat hcihw ssecorp ehT :ssecorP lamrehtosI .dnuopmoc cinagro nevig ni negortin fo ecneserp eht smrifnoc noitaroloc dutg ro eulb naissurp fo noitamrof ehT eulb naissurpÂ ÂÂÂ ÃãÂ
ÂÂÂΓÂΓÂEτÂΓÂEEEEE(EÂEEEEEEEE(ÂEΓÂΓΓÂΓÂ(ÂΓÂΓÂEΓEΓÂΓEEEEEE³ ÂÂ ÂE³EτÂ ÂτÂ ÂλÂ ÂλÂ ÂτÂ ÂλÂ ÂτÂ ÂτÂ ÂτÂ ÂτÂ ÂτÂ Â ÂTOΓÂ ÂτÂ Â Â ÂTaN21 + 3]6)NC(eF[4eF Â Â Â Â ÂTΓ3lCeF4 +]6)NC(eF[4aNÂ ÄRIGHTτd ³Â ÂτÂ Â ÂτÂ Â ÂτÂ Â ÂλÂ Â ÂλÂ Â ÂτÂ Â ÂλÂ Âτ .lCH detartnecnoc fo ssecxe dna noitulos edirolhc cirref elttil a dedda
neht si noitulos delooooc eht oT HOaN2 +]6)NC(eF[4aN Âτ¢ 2)HO(eF + NCaN6Â Â TτÂ Â ÂτÂ Â Â ÂΓÂ Â ÂΓÂEEEE³EEEEEEEEEEA³I3ÂI3I34 I34 .taeh fo tnuoma egral esaeler ria ni snrub tI :riA ni nruB .desu stnatcaer fo ytitnauq eht nopu sdneped noitcaer fo yplahtne ni egnahc ehT .HÂτÂibb detoned si tI .laC K 4.62 ÂT¢I = 1HÂTOΓ)g(OC ÂTO 2/1
+)s(C :edixonoM nobraC aiV 2OC fo noitcaerF .iii tnatsnoc lleC .retemirolac dellac si sutarappa elbatius a ni derusaem si noitcaer lacimehc a ni devlovni taeH bR/a2 = iT .c ssecorP citabaidA .2 noitamroF fo yplahtnE fo noitaluclaC .ydob eht yb debrosba si ygrene eht ,deppord si erutarepmet nehw dna itaidar citengamortcele stime ti ,detaeh si ydob a
nehW tropsnarT lacirtcelE 1:repaP muirbiluqE esahP 1:repaP .elbissop ton si noitanimreted latnemirepxe hcihw To this sufficient amount of silver nitrate is added when a white precipitate of silver salt is obtained. Types of reagents: electrophilic and nucleophilic. Viscosity c. isoelectronic ion with benzene. Hydrolysis: In hydrolysis, diborane forms
boric acid. b. Unit Cell: The smallest portion of a crystalline lattice is called a Unit Cell. Thermochemistry: Exergenic and endergenic compounds, reaction enthalpy at constant volume and pressure, combustion enthalpy, bomb calorimeter, neutralization enthalpy and ionization enthalpy, Kirchoff’s law, Hess’s law, dissection energy association of
connections. Formation of CO2 directly: C (s) + O2 (g) – CO2 (g) -SH = – 94 K Cal. of BaSO4. When the borazine is pyrolysis above 340 °C, B3N6H10 and B5N5H8 are produced. Some important applications are given below: 1. The precipitate is separated by filtration, washed successively with water, alcohol and ether and dried in the steam oven. The
HCl ¢ V ml. According to the above law, the interceptions of any face, like KLM, on the same three axes will be multiple simple integers of a, b and c, respectively. It is easily determined by dividing the molecular mass by the density of the compound (i.e., V = M/D). Here, r+ is the radius of the cition and r- is the radius of the surrounding anions.
Whereas, in the formation of SnCl4, the metal âÂ´Sn' used both âÂ´5s' and âÂ´5p´ electrons. Indexes Millar b. Diborane and superior boranes. The reluctance of two ns-electrons to participate in the formation of bonds due to the high energy required to unpair them is called the inert pair effect. The enthalpy of the reaction depends on the temperature
and pressure of the reaction. Why is the transport number of the Li+ ion smaller than that of Na+? Properties that depend on the amount of substance present in the system are called extensive properties. It is a gold test for the detection of nitrogen in all classes of nitrogen compounds. Aq+ (aq) +)qa( )qa(The heat of neutralisation is the heat
produced when one mole of water is formed from the reaction between an acid and an alkali. Let us consider 'n' moles of an ideal gas enclosed in a cylinder fitted with a frictionless, weightless and movable piston. Generally, the melting and boiling point of elements increase down the group. Peroxydisulfuric acid is the inorganic compound with the
chemical formula H2S2O8 also called Marshall's acid after its inventor Professor Hugh Marshall. [Hints: ÃÂx = ¢ÃÂÂ/2.ÃÂp ÃÂx = 5.27 ¢ÃÂÂ 10-35/3.3 ¢ÃÂÂ 10-2 1.59 ¢ÃÂÂ 10-33m] Crystal Lattice: Crystal Lattice is a three-dimensional representation of constituent particles (atoms, molecules or ions) arranged in a specific order. Chemical
Propeeties: 1. Reaction with Ammonia: Diborane react with ammonia but the product is depend upon the experimental condition. Simple cubic b. Water and sulphur system, two component system Ag-Pb, KI + water, eutectic point, Formation of compounds with congruent melting points, Deliquescence, efflorescence, triple points. There are total 14
possible three-dimensional lattices. Write short notes on the followings: a. ¢ÃÂÂ By joining the lattice points with straight lines the geometry of the crystal lattice is formed. Hess' Law of Constant Heat Summation: Hess Law of Constant Heat Summation is also known as Second Law of Thermochemistry.
Ã ÂB2H6 + 6NH3 -----> 2B3N3H6 + 12H2 Properties: 1. In increasig order of stability: CH3CH2CH2+, CH3CH+CH3, (CH3)3C+ b. This is law of corresponding states. So, down the group the carbonates become more thermally stable. Preparation and properties of aliphatic amines. Discuss i.
Separation and identification of 1Ã°Â , 2Ã°Â and 3Ã°Â amines. Internal energy and Enthalpy b. Define i. V volume of N HCl is required for complete neutralization of NH3 evolved. So, Neutralisation is an exothermic reaction and always produces heat. Basic idea about Dimethylzinc, dimethylcadmium, alkyl and tertium dialkylan. Name the alkali
metals and give your electrical configuration. Conditions constant c. Mechanism of organic reactions: homoplass and heterology of covalent titles. This use is supplied by its unique molecular structure, antioxidant effect and biological compatibility. A yellow or PPT color. Compounds of organosulphur: Preliminary Idea of Organic Sulfomides,
Sulphones and Sulphones, all of formation and chemical reactions of the Toles and Thioeta. State Vander Waal Equation and explain the meaning of Vander Waal constant. Ã, å "e.g. "Q = the isocoretic process: the process that occurs in the constant volume is called a isocoranic process. Draw molecular orbital power diagram for oxygen molemplate.
Chemical connection: COVALENT CONNECTION: VB theory and its limitations, directional characteristics of covalent venue, hybridization and form of inorganic moleps and are vsepr Theory with special reference to the venue and electronegativity, moethine. A deep violet color Indicates the presence of sulfur. In this cyclic compound, three BH units
and three NH units are alternated for each - another. CCL4 D. 9. Therefore, = (8 x 1/8) + 1 = 2 So, the number of particles per unit unit unit for the BCC unit CÃ © 2. Discuss the BP variation of H2O, H2S, H2SE and H2TE c. If sulfur It is present in the organic compound by means of a fuss with Na Reaction to form sodium sulfide. Organize the
following a. KOH is a stronger base than BA (OH) 2 II. Therefore, only 1/8 of the constituent particles, belong to a particular unit CÃ ©. Thermodynamic: Thermodynamic Objectives, Thermodynamic Terms, First Law and Your Mathematics, Internal Energy, Enthalpy, CP and CV Relationship, Joule - Effect Thomson, Joule Thomson Coefficient to ideal
a real gases, inversion temperature, work done in irreversible , Reverse and onobrac od ritrap a m©Ãbmat e onobrac od ritrap a etnematerid odamrof res edop onobrac ed odix³Äid O Â Â¢Ã.sapate sair¡Âv me uo apate amu me arroco oÉÃ§Ãaer a reuq ,amsem a ©Ã acimÂuq oÉÃ§Ãaer amu ahnapmoca euq latot rolac ed oÉÃ§Ãairav a euQÂÂ¢Ã
,ecelebatse iel A .aiplatnE dnoB ed oÉÃ§Ãaicossid a euqilpxE .3+ lev¡Âtse siam arap 5+ ed oÉÃ§Ãadixo ed odatse ues adum e ,arotuder aicn¢Âtsbus artuo a etnemlicaf uodixO .sosorret sonilacla siatem ed sodamahc oÉÃs ossi roP .ocir¡Ãbosi ossecorp ed odamahc ©Ã etnatsnoc oÉÃsserp a erroco euq ossecorp O :ocir¡ÃbosI ossecorP O=VÂ«ã .e.i Ã Ã Ã
Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã .b orgen oproc od oÉÃ§ÃaidaR .oÉÃ§Ãacilpxe ed esaf ed arger e odatsE .4 lCsC ed anilatsirc aruturtse a atucsiD .etreni rap ed otiefe mebixe .cte) iB (otumsib ,)bP(obmuhc ,) nS (ohnatse ,)oP(oin´Âlop ,)lT(oilºÃt omoc ,p ocolb od sotnemele snuglA .b socilÃtem-N sodavireD .opurg on atnemua)3OCM(sosorret sonilacla siatem ed
sotanobrac sod acimr©Ãt edadilibatse A sÃ Â¢ÃfrottiH rop oÉÃ§Ãanimreted aus e etropsnart ed oremºÃn ,sÃ Â¢ÃhcsuarlhoK ed iel ,snoÃ ed oÉÃ§Ãargim ,acin´Âi edadilibom ,aicn¢Âtudnoc an oÉÃ§Ãiulid ad otiefe ,ralulec etnatsnoc ,ralom e acifÃcepse ,etnelaviuqe aicn¢ÂtudnoC ,sacitÃlortele seµÃ§Ãulos me aicn¢ÂtudnoC :ocirt©Ãle etropsnarT
.etreni rap ed otiefe oa odived aluc©Ãlom ed oÉÃ§Ãamrof arap levÃnopsid ¡Âtse oÉÃn 'nS' ed nort©Ãle 's5' ,saM .edaditned aus medrep socim´Âta siatibro so ,seralucelom siatibro ed oÉÃ§Ãamrof a s³ÃpA .oterp tpp mU .saluc©Ãlom san setneserp soelc©Ãn so sodot a ecnetrep ralucelom latibro mu me nort©Ãle adaC .zul ad edadicolev ad %05 ed
edadicolev amu moc odnajaiv)g82-01 x 11,9 = m(nort©Ãle mu a odaicossa adno ed otnemirpmoc o raluclaC .laedi otnematropmoc od laer s¡Ãg od oivsed o reverceD .b 5FA e 4FS ,3FrB ,2-3OC ,2FeX :sotnemele setniuges sod oÉÃ§Ãadirbih a e aruturtse a ,amrof a racidnI .orgen oproc od oÉÃ§Ãaidar ad lartcepse oÉÃ§Ãiubirtsid a acilpxe airoet atsE
.adarutasni oÉÃ§Ãaer amu artsom elE V .laedi s¡Ãg mu ed acit¡Ãbaida oÉÃsnapxe .onobrac .onobrac ed odix³Ãnom od of crystalline networks: Ã ¢ Â Ã ¢> Each constituent particle is represented by a point in a crystalline network. It is used as an anti-aging and antidane agent in the cosmetic sector. 2 "H = 3" H1 + 3Â "â € œ" "H2 participate in the
lessons live using any device, whether it phone, tablet or computer Start today better * GrÃ TIS * chemical prof ever. .No only the chemical notes are organized and easy to digest, but also author has many examples and tricks. The compound is isoelectrane and isostructural with benzene. Weak interactions: H-binding forces and van der Waals. What
is the difference between âƒ and â € bound. Interplanry spacing in a cymbian system, ray and number of coordination, packaging of octaic and tetraÃ © dricos. BoyleÃ ¢ â € â € ¢ temperature d. The faces of the crystals and also the plans within the crystals can be characterized by means of an adequate set of coordinates. Troutton â € ¢ rule b. The
state of oxidation of the sulfur is +6. The molecular orbital wave function (Â Â¨) is such that Î '2 represents the probability density or trone load density. When a gas is allowed to expand from high to low pressure through a porous outlet on adiabatic conditions, the gains is cooled. At the ourn at the À to the om to the à the om to the Ã é à Ã 4bf3 +
3nabh4 = 2B2H6 + 3NABF4 Properties: Properties Physics: * It is sensitive to air, volatile and gas reactive with repellent smell. Or the geometric arrangement of constituent particles of crystalline solids as spot in the space is called crystalline structure. # Effect of inert pair, affect the fusion point and boil of elements. Â "H = Ã ¢ â € â € œ94 K Cal.
These distances in interceptory calls. Repeating in different directions, CÃ © Lula Unit generates the entire network. If two liquids are taken with the same superficial tension, whose molecular masses are M1 and M2 and whose densities are D1 and D2, respectively, then, then, "m1." = moc moc sodiuqÃl siod ed sadeuqarap ed oÉÃzar a ,missA
)2D/2M(/)1D/1M(= tser eht DNA metsys dellac si si yduts rednu hcihw esrevinu trap under eht .negyxo thing dezidixo too tnrub yletelpmoc she ecnatsbus under elom 1 nehw egnahcxe taeh eht in denifed ecnatsbus she was under noitsubmoc under yplahtne ehT eulaV cifirolaC under noitaluclaC .noitauqE USA ¢ Garb Si Tahw .b Elpicken Yb
DNAUDHY HGIA 3FETY HG 3 + EB Una ytilibulos hgih ehT .c 3OCrS DNA 3OCaC, 3OCgM, 3OCeB under ytilibats lamreht under noitairav eht ssucsiD .scillatemonagro Lyra DNA lykla noitcudortni the OT, seicnednet gnixelpmoc DNA noitavlas, sedirdyh, pihsnoitaler lanogaid, yduts evitarapmoC: stnemelE kcolB-and .latibro ralucelom under eht
EPAHS EHT DNA YGRE EHT EHT Enimreted HCIHUN METNAUF Metibro RaAe HCAAF HCAAF HCAAF HCAAF HCAEMUSA SNOF. YHW, 6fo DNA 4FO MRO SNAGE 0FS DNA 4FS TEHEA CRAUEHE Â ã, â ã,, t t,, t t t t t t t The Century the '€ Y clans wereh.f yb htorf tub yroeht snht snOirep 1: erofrepmet. B Walk S-€ HCMULTHLOK EHPMOCKS
ESOPMOEQ EHT SNH LACHT LACHT LACHT LACHT LACHT LACHOQE Eud #: Securellower .slatibro Rajrom under Noitamrof Eht Owtc.van ictang emosel .Vilaks KCAB YB dezilartuen dicken via tnuomma eht g nidnif yb denimreted she ainomma under tnuomma ehT .semulov ralom under oitar eht OT lauqe she noisnet ecafrus emas of the universe is
called around. Thus, CL + on â € NaCl â € â € œNO, â € I + in â € nai acidify a pity of hyno extract with HNO3 diluted and adding agno3 solution. Mâ® "V \ ¥ ¥" / 2 to x. The excess of Ammonia is boiled. Indicates the presence of fulsories. It is denoted by E. Phases II. As the ideal gas does not have such strengths, therefore, there are no energy
expenses to overcome these forces during the expansion. As a part of it is used to overcome the force of vaner waal at the expansion between them during the expansion. When a light ray falls into a crystal plane composed of regularly organized particles, the radiographs are diffrescribed. (FCC) Unit Cell: In the CÃ © Lula Cábria Cubica Centered on
Face, the Particles are present in the corners as well as in the center of the face that is shared between two adjacent particles. What is your use? According to the KOPP rule, it was found that the molar volumes of two members of a homophage homophage of organic liquids differ in about 22 ml, for each CH2 KOPP group calculated the volume
equivalent of each element Middle of a simple medium arithmetic. WT of salt AG Â «2) x 108 + 2 = 140 g The substance given is mixed with double the amount of sodalime and heated in a test tube. How much can be taken as behaving as a stem of particles with mass, energy and impulse. Fullerene has incredible electrical conductivity and has a
stronger conductivity than copper. Explain the first thermodynamic law and its limitations. Substitution) Ã ¢ € = (14/1000) .vn (100 / x) = 1.4VN / x where: n = normality and v = volume of the acid used. R-X + MG - R-MG-X Some important reactions: Paper: 2 Organo sulfur compounds Paper: 2 villains and ketones Paper: 2 Organic compounds of
chemical nitrogen physics .rolav ues avercsE .sordeasoci rezaf arap sarger moc odroca ed sodÃurtsnoc ,sianogaxeh e sianogatnep si©Ãna me adaesab aruturtse aus ,ritsixe edop sonerelluf ed otinifni oremºÃn mu ,etnemacireoT .tw/001(otoza ed .adagujnoc esab a azilibatse siam e omsem is a oÉÃ§Ãerid me snort©Ãle so iarta siam ,lartnec omot¡Ã od
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ronem ©Ã 'oP' ed oÉÃ§Ãilube e oÉÃsuf ed otnop o ,olpmexe roP .erocarap omret o euqilpxE .sasnetxe sedadeirporp ed olpmexe o oÉÃs cte ervil aigrenE ,aiportnE ,aiplatnE ,selom ed oremºÃN ,emuloV ,saM .2sn ©Ã aicnªÃlav ed ahcnoc ed oÉÃ§Ãarugifnoc auS .aR e aB ,rS ,aC ,gM ,eB oÉÃs selE .otreba ametsis ed odamahc ©Ã etneibma o moc
air©Ãtam e rolac racort edop euq ametsis O .3OS2H me 4+ e 4OS2H me 6+ ©Ã S ed oÉÃ§Ãadixo ed odatse O .acil¡Ãtem atarp ed oudÃser mu raxied arap meµÃpmoced es ,sodiceuqa meres oa ,euq ,siev©Ãlosni atarp ed sias meramrof ed otaf on es-aiesab odot©Ãm o oÉÃ§Ãanimreted e oÉÃ§ÃamitsE 2:lepaP 2H21 +6H3N3B2>â.T.Hâ 3HN6 + 6H2B3 Ã
Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã 2H6 + NB2>â.T.Hâ 3HN2 + 6H2B Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã 3HN2.6H2B >â.T.Lâ 3HN2 + 6H2B Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã ,IdeIlnis YX2 = Ã Ã Ã Ã Ã Ã Ã Ã Ã ,IyTn ,otnatrop ,«ânis YX = Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã ZY + YW = ohnimac ed aŞÃnerefiD ,missA ÃÃ Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã
Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã Ã .ii aN euq od roiam ©Ã iL ed oÉÃ§Ãazinoi ed laicnetop O .oditrever res edop oÉÃn e apate amu me erroco euq ossecorp O :levÃsreverri ossecorP .tw = otoza ed % mg N .adaraperp etnemetnecer oid³Ãs ed odisurportin ed oÉÃ§Ãulos amu oÉÃ§Ãrop amu a ranoicidA .1 Beso4 and MgSO4 are highly sunny, Case4 is moderately
sunny, but Baso4 is practically sunless. Each carbon atom is attached to other metals and is hybridized SP2. The most stable and common fulleriche is © the molliacula C60. A piece of paper that is a metal metal - 2Âº 1. Borazine is slowly hydrolyzed by Water to produce rich acid, ³ and hydrogen. The hydro³lysis is favored by the increase in
temperature. Explain the effect of dilution on molar conductivity, equivalent and specificity. Energy energy and coupling of connecting bonds: Whenever a bond is formed, the energy is absorbed. Let the press of the g and P-DP be the external press under which the volume of the g increased by DV, then work carried out in this expansion. CO2 and
NO2 II. Moreover, ductility is 100 times stronger than the nation. Rmica and endot ©rmica II. The standard enthalpy of combustion is the change of enthalpy when a mole of a substance burns under standard state µ. Any solution that is transferred in an excess BAKER and BACL2 solution is added to it for full PPT. Dipole moment: Percentage ionic
character in HX, molecular geometry of polyatomic molecules. When X-rays fall into a crystal, they are scattered. Ic ic Â· ie ie ie = Ã¿Â½ reversible process: the process that occurs infinitesimally slowly and whose directionÃ¿Ã¿Â ⊆ o at any point can be reversed by applying an alterÃ¿Ã¿Â ⊆ o infinitesimal in the system state called the reversible
process. ~> For Unit Cell: In the body-centered cubic unit squid, the particles are present in the corners as well as in the center of the body. White Ppt. The amount of heat needed to change its temperature to a degree of a subst constant press. Compton in 1923 noted that the duration of the wave of radiation od od adno ed otnemirpmoc o euq roiam
erpmes ©Ã)' ®Ã(Radiation () I.E. "> The change in the wavelength is independent of the wavelength of the incident radiation, as well as the spreader. Principle of the uncertainty of Heisenberg: It is not possible to determine with precision and simultaneously the pulse and the position of small particles on the move. For example, the +5 oxidation
state of +5 is less stable than the +3 oxidation state. Determin Of the molecular mass of a healing cord by silver salt and an organic base per all of the chloroplatinate salt. Â € B2H6 + 3O2 = B2O3 + 3 H2O 5 . CO (g) + 1/2 O2 (g) â € CO2 (g) â € "H2 = Ã ¢ â €" 67.6 K Cal. Fullerenes are used â € < Antivirals. If the waves are in phase after reflection,
the difference in the distance traveled by the two rays should be orqaul for an integral number of wavelength, NÃ Â 'for constructive interference. Therefore, Z = 8 x 1/8 = 1 So, the number of particles per single cell phone units for simple mobile It's 1. Broglie hipotesis c. Leave ox, oy and oz Represent the three crystallic axes and let the ABC be a
unit plan. Hydretes of boron. Effect of inert III. When the number of reagent moles is duplicate, the enthalpy change also becomes pair. The preparation: 1. When an acid and alkaline reacting, the heat is distributed. A mixture of acidic acid and acyptic anhydride b. Exotious and endottion 2.Amic reaction. Discover an expression for the work done in a
reversible expansion isotemic b. Why the CO2 is the Gás, but the SiO2 is solid? On a crystal truss, all corners are shared by eight adjacent unit skills. In growing order of agriculture: HCOOH, CH3COOH, CH3CH2COOH C. The PPT. Once again, a GM of organic compost contain (32x b) / 233 GM S. nine issues to be defined. Ã, å "e.g. "T = the adiabal
process: the process in which no heat changes occurs is called adiabaly process. Explain i. Derivate the laeR laeR uo rednaV ed LAAW ed equation. Personally, I liked two things in these notes: a) The gradual depth of concepts, which saved me the overwhelming sensation that feels when (disorganized) calls of the so-called "obligatory didothic books"
comes exploding from all the places; B) The abundance of problems and their detailed solues! In addition, it seems strange, but the teacher - on your site - promises to respond freely and instantly any email asking a chemistry question through your website. YÃƒÂ ± Ã¡nénica 2. The first law does not specify this process is viable or not. ~> Trellis points
on a crystal trellis are united by straight lines. Entropia 5. Kohlrausch Law II. À, angle between border B and C is A, A and C and between A and B is a. -% sulfur in organic compost = (32 xBX 100) / 233 x A paper: 2 Structure and Control Paper: 2 Organic Reaction Mechanism Paper: 2 Organometable Compounds MÃ © All of preparation: the grignard
reagent can be prepared by the alkyl halide reaction with metal mg in presents of it is dry. * It is quite stable in the absence of moisture and grease. The fulerene molemplate can be used as an antioxidant because it can easily react with radicals due to high affinity of the electric elé. Solid Ammonia Indicates PPT of yellowish chlorine. Each fan motor
performs a discreet energy that is directly proportional to the wave-length frequency. Calculate the ch3cooh molar conductive in the infinite dilution. That among you (22), V (23), CR (24) and MN (25) has the maximum second ionization energy. Ã, â € B2H6 + 6CL2 = 2BCL3 + 6HCL 6. If the heat is evolved, the reaction is said to be exotic and when
the heat is absorbed the reaction is an endotan RMICO. Hot water in a temporal bottle is an example of this system. After counterfeiting with herself, halogens in the organic compound are converted to the corresponding sodium methods. The geodemen and electronic factors of connection in the dna emulov ralom gniylpitlum yb deniatbo tnatsnoc a si
rohcoehR .slatibro cimota eht llif yeht sa yaw emas eht ni slatibro ralucelom eht llif nortcelE .deniatbo si elozarob detaeh era 3:1 fo oitar eht ni 3HN dna 6H2B nehW:noitaraperP .noitcaer cimrehtodne na rof evitisop dna noitcaer cimrehtoxe f evitagen si HÂÂ stnatcaeRHÂ Ã sesaerced setahplus)AII(latem htrae enilakla fo ytilibulos ehT .metsys
eht fo ygrene lanretni rieht dna selom fo rebmun eht si os dna degnahc osla si emulov eht ,degnahc si sag eht fo ssam fI .noitulid etinifni ta ylevitcepser 1-lom2mcs 0.19 dna 9.524 ,4.621 era etateca muidos dna dna dna dna l1 CaN fo seitivitcudnoc ralom ehT .2 emalf ni roloc wolley nedlog sevig aN .)0001/41(= 3HN fo lom V ni tneserp N fo tnuommA
N mg 41 ro 3HN fo mg 71 sniatnoc 3HN N lm0001 3HN N .b tnatsnoc cificepS .a ni llec tinu rep smota fo rebmun eht etalucla? stnatsnoc lacitirc si tahW .senerelluf ,sedirdyh orob ,enizaroB .elpmaxe eviG .B tip cirolhcordyH .metsys detailsI dellac si gnidnuorrus eht htiw rettam ron ygrene rehtien egnahcxe nac hcihw metsys ehT :metsyS detailsI .c
requestI negordyH .gnoleb yeht elbat cidoirep eht fo kcolb hcihw o)85(eC dna)97(gA ,)05(nS ,)21(gM -elbat cidoirep eht ni stnemele gniwollof eht fo doirep dna puorg eht yfitnedI .smota suoesag detailosi morf demrof era sdnob fo elom 1 nehw desaeler ygrene fo tnuoma ehT ,sa denifed eb yam dnoB .snoitseuq ni dexim eb yam stinu tinu nereffid dna
noitseuq a ni strap lareves eb yam erehT .enirolhc htiw tcaer yltneloiv ti :enirolhc htiw noitcaeR .i yb naem uoy od tahW .5 ecnatcudnoc cificepS .S mg 23 sniatnoc 4OSaB mg 332 ÂµÂÂ¢ 332 = 4OSaB fo thgiew ralucelom :noitaluclaC mg b fo .noitulos etateca dael da neht dna tip citeca htiw tcartxe eht fo noitrop dnoces a yfidicA .elucelom eht fo
ytilibats eht rof tnuocca tnuocca Root of co-efficient viscosity. It is a state function depends only on the initial and final state of the system. Neutralization enthalpy is always constant for a strong acid and a strong base: this is because all strong acids and strong bases are completely ionized in diluted solution. Thus, with two stage f shape the Bef2
molemple. or c.c .. Calculate the position of uncertainty. Ã, Å "The limited radio rate is the minimum value allowed for the ratio of Radii Catii Anionic Cathi (q) for the structure to be stable. Block Elements P: Comparative study, relationship between metal elements, non-metal and metalloids of group 13-17, elemental idea of hydrides, oxids and
halides. Discover an expression for the work done in reversible expansion isotemic. Cálco: Let the weight of the organic substance be X GM. Where, n = 1,2,3 ... (difficulty order) "= X-ray wavelength" D = distance between airplanes ã, angle, in which interference occurs . Fullerene is more difficult than the diamond. Why the vapor pressure of the
pure solvent is more than that of the solution? Methyl derivatives 4. MacLeod in 1923 gave the following relationship between the surface tension (Y) and density (d) for a normal-c = ³1 / 4/4 / (d ICD) where d is Liquid steam density at a certain EC temperature is constant. Solubilities of alkaline earth metal sulphates decrease in descending the group.
The temperature of the water is observed before the substance is inflamed by a high chain. 8. As Glycerol reacts with the following: a. 2NA + S â € Na2s Thus, the obtained sodium extract of the fused mass can be tested as: a. which between H2SO3 and H2SO4 are more agricultural? These two orbitals (ie 2s and 2P) are hybridized and form two HB
orbitals and organized linearly in 3D space. For this reason It is sometimes referred to as â € œGenzene inorganic. By defining the questions, the entire program can be addressed as much as possible. 4. Ã ¢ â ~ Â Â ¢> Skills of simple czbic units: On the simple cyt unit cell, the particles are present only in the corners. It is a colorless liquid 2. The
carbonates of the IIA group thermally decompose and produce methyl and carbon dioxide. This means that the universe is the combination of system and surroundings. For an infinite volume change of V1 for V2, the total work done during expansion ivors at the ivors to the ivors , Ivors, ã, ã, ã, ã, to, where P1E P2SÃ £ o, respectively, the initial
pressure and the final press. Procedure: The unknown acid is dissolved in water and treated with a slight excess of ammanship hydroxide. Bravias Lattices II. The NH3 Q. Fullere weighs only 1/6 of the copper material. 3. Calculate for the number six. According to this principle, energy can not be created or destroyed, can only be transferred or
changed in one way to another. SNCL2 is more stable than SNCL4 because in the formation of SNCL2, the metal 'sn' uses only two trons '5p'. What is conductivity? They are. Two months of methyl derivatives. PCL3 then, PR = P / PC or, p = PRPC â € ™ â € (equation 1) vr = v / vc or, v = vrvcan, ã, , Â € (equation 2) in a similar way, tr = t / tc or, t =
ttcan, ã, â € (equation 3) now replacing the value p, v and T In the equation of Vander Waal, we obtained- (PRPC + A / VR2VC2) (VRVC Â Â Â'B) = RTRTC ã, â € (equation 4) placing now the values of the chronic constants in equation (4) we get- (PR A / 27B2 + A / VR29B2) (VR3B Â Â Â Â Â Â Â Â Â Â Â Â Â this equation is completely free of constants
like R, A, B, so this is applicable to all n fluid state. Give a whole to your determination. The amount of heat needed to change its temperature to a degree of a substance. Internal energy and enthalpy e. Closed system: the system that can only exchange heat, but does not matter with the environment called closed system. It denoted by the letter
Capital 'R'. During the neutralization phase, the hydro-ion anions of Ahcid reacted with hydro-³ ions of Ahlcalis to form Water. C6H12 C. degree of freedom III.number of the component with appropriate example b. Each associated wavelength with a defined energy value and the total energy of the molecule calculates the sum of the energies of the
occupied molecular orbitals. Structure and glue: hybridization and geometry of lengths of hydrocarbon bonds, angles of bonds, energy of dissociation of bonds, localized and delocalized chemical bonding, Vander WAAL interactions µand resonance of hydrogen bonding, hyperconjugation, inductive and electron effect, their effects on the properties of
compounds. What A© azeotropic mixture? The magnitude of the trellis energy remains almost constant because the sulfate is not large. The energy of the radiance that is not © - e = HÃ®id (where 'H' is the surfboard constant and 'Â½' The frequency of radiance) or, and - The neutralizes the area between a acid and a base to form a salt and water.
Sulfrich acid C. As a result, the '6S'-elÃ ©tron de' po 'I'm less available than' 5s'-electron If a body absorbs all the radiation µes that fall on it, there is a call of a black body and the radiation emitted by it. Now, 'Po'-Element contains the maximum number of el ©trons' D 'e' f '. Write down the structure of peroxodisulfRich acid 5. The following points
are the main points of this theory - 1. Reaction with concentrated solution of KOH: with concentrated KOH solution, forms the potassium hypoborite. C2H5OH (L) + 3O2 (g) Â2CO2 + 3H2O ~ Â Â298 = 1366.8kj. For elements. elements. In different alterations µaltrac³s, the heat of the different area is different if different altr³hpic shape is involved in
the area. How a primal, secondary and third anima- tion reacts with nitrous acid provides ³ of discretion in the dialkilo of Latvian. As a result, the C60 behaves like an alkene deficient in elÃ ©trons and reacts easily with sppA ©cies rich in elÃ ©trons. UnitSquidCharacter: ~> A unit squid has three edges A, B, and C and has three angles a, Â³, and y
between the respective edges. Carboxacids: They are all general preparedness µ and properties of monocarboxylic acids and their derivatives, such as Ã© ster, acid chlorides, amides and anhydrides. Vander Waal State Equation and explain the significance of the Vander Waal constants A and B. The drop in temperature (DT) produced by the drop in
pressure (DP) under the additional condition J-T effect call. The resulting two effects I.E. "Hydrate" HLATTICE - "Hydrate" becomes more negative As we move from being (oh) 2 to BA (OH) 2, it explains the increased solubility. Inter halog compounds At â¬ tb = a / rb we know this- (p + a / v2) (v Â"b) = rt Ã¢ (for a mole) or, PV 'BP + A / V' AB / V2 =
RT by multiplying the above equation by v2 / p, we obtained v3 Â"V2B + AV / P Â° aab / p @rtv2 / p = 0 or, v3 '(B + RT / P) V2 + AV / P' AB / P = 0 when, t = tc and p = pc Â ³ v3 ' B + RTC / PC) v2 + AV / PC 'AB / PC = 0 —(Equation 3) 3VC2 = A / PC now putting VC and PC value into (3) we get -9b Ã ÂÂÂ b = 27b2RTc/a or, 8a = 27bRTc or, Ã
Tc = 8a/27RbÃo Determination of van der waal constants, law of the corresponding states, reduced equations of the states, Boyle temperature Âs. This vessel is equipped with a stirrer and a term with a reading up to 1/100 degree. OnÂ For end-centered (CEC) composite unit: In the face-centered cubic unit squid, the pieces are present in the corners,
as well as in the center of two faces that are opposite one side to the other. another. Weigh about 0,2 g of dry silver salt into a crucible and ignite it until the decomposition is complete. lead sulfide confirms the presence of sulfur. This was proposed by Sidgwick. It was introduced by Newton Friend in 1943. This is called the law of rational µ or
indexes. Explain two of the following µ: 2. Organometallic Compounds: Organomagn compounds ©sio:The shape structure of the Grignard reagent and the application in organic synthesis. The temperature drop is due to decay of the acidic energy of the molecules of g. Ag salt/wt. Be + 2NaOH â> Na2BeO2 + H2 Be displaces the hydrogen sulfide
acid hydrogenated dilute, but with hot concentrated sulfuric acid, releases the gÃ¡s SO2 Be + H2SO4 â µ ³ ³ ³s Trillise dÃohydro 5. Organic nitrogen compounds: Classification, nomenclature and structure of amines. A 100 g of organic compound contains ©m (32 X b X 100)/233 X a gm S. The ignition is repeated until the crucible containing the
silver residue has reached a constant weight. According to Avogadro's hip³thesis, the molar volume of all gases in S.T.P. is 22.4L. How's propene glycerol synthesis going? C6H6 B. Minor Cátion, the greater the ability to polarize Ã ¢ nion. Uses: One of the owT 7.95 = 9.91 X 3 = sdnob bud erhT 6.93 = 8.91 X 2 = O2 8.26 = 7.51 X 4 = H4 6.15 = 6.8 X
6 = C6 2-erutcurtS 532 = rohcaraP detaluclaC; S 4.1 = 4.1 X 1 = gnir derebmem 6 = 6.97 = 9.91 X 4 = sdnob bud ruf 6.93 = 6.8 X 2.9 = 4 X 2.9 = 28 X 4 = 2.7 = 2.5 6.15 = 6.8 X 6 = 6.8 X 6 = C6 1-erutcurtS -ENONIUQ: desoporp erutcurts owT:elpmaxE.mota-B fo slatibro-p ytpme dna mota-N fo slatibro-p dellif ehneewteb gnippalrevo esiwedis eht morf sesira
hcihw ,dnob evitaD a si enizarob dnob (N-B) .latsyrc diuqil ,sretemarap ybulos ,erusserpretniNocNoitchNoitchNkNoitchR, erohcaraP, emulov, raloM, stnemerussem, rihat, ytsocsiv, noisnet, ecafruS, elur, snotuorT, erusserp, ruopav, diuqil, fo emulov, eerF: etatS diuqiL .c cibuc, dertnec, ecaF .asrev-eciv, nwonknu si mutnemom neht nwonk, elcitrap and
fo noitisop fI .a :sgnilof ehalpnixE .metxeoEoReeo, eretsufEreereeereeeo htin sil .secittaL siavarB3yb nwonk6aCL connection to walkC.1 dn2-repaP2rtsimaC lacisyhP.(E) mtsisAh3fo ygrene ylpmis ro ygrene lanretni dellacSi mtsis a htiw detaicossa ygrene fo smrovsmrA.c enezneborolhcid-penezneborolhcid-o .seitrpDNAlacisyhp,noitaraperp fo sdohteM
:LaDyhonoEmcncncncnostiNofsvcncncncl:NofsvcnostalNcnofsvclANcnofsvcncncncncncl:Ncnophaphacl C dellac si noitaidar derettacs ehf fhtgnaw ni egnac ehT .oitar suidar gnitimil yb nnm uoy od tahW .enerelluf retsnimkcub si enerelluF fo eman laer ehT .ytgrip evitutitsnoc na evitidda na htop si rohcaraP .tw mg a = dnuopmoc cinagna fo
.6H6C2CNgnnob echtiw selucelyoo selenagro deseneoruSeiatseftpsiRetsefsiRetcncncrRec-RetserepNcNc rO -3ON fo arutcurts gnitanoser htirW .tri nniemar snortsele-'s' owesahat dna noitcaer lacimehc no etapicitrap ot ycnednet ssel a wohs snortcele-'s' owt eseht ,yltneuqesnoC .c .enicidem si enerelluF fo sesu tnatropmi Ring medus = 2 x 1.4 = 2.8
thus, calculated parachor = 216.5 since the experimental parachet Ã© 236.8. Therefore, the structure - 1 is correct for quinone. Hund and R.S. Mulliken in 1930 and later developed by I.f. Lennard Jones and Charles Coulson. Diatoms homonuclear and heteronuclear [CO, NO], binding force, binding energy. The resulting liquid is then treated with
excess alkali and the released ammonia is absorbed in excess standard acid. These products are analogs of boron-nitrog of diphenyl and naphthalene, respectively. However, not much use has been shown in solving structural problems. Reaction with dry HCl: Reacts with dry HCl in the Al2x6 pressence, formulations B2H5Cl Ã¢ â ¬ B2H6 + HCL =
B2H5Cl + H2 2. That's why ³ ³ Cao A â¬ b3n3h6 + 3hcl —> b3n3h9cl3 Ã Ã¢ âTM CH3C (NH) NH2, CH3CH2NH2 , (CH3) 2nh, ch3conh2 d. Explain the following: a. Reactive intermediaries: carbanÃs, carmouthµÃes and free radicals (generation, structure and stability). Because of this, BIF5 acts as a strong oxidizing agent. CAA + 2HCL —> CACL2 +
H2O ZNO + 2HCL —> ZNCL2 + H2O ZNO + 2naoh —> Na2ZNO2 + H2O 4BE: 1S22S2 in the state of the soil, has no unpaired elecron But in the excited state has two unpaired elecrons. Paper: 1 block elements group 2 elements of the peri³hadic table are called alkaline Earth metals. Thus, the polarizing capacity of the IIs ion group decreases the
group as the size of the cap increases the group. Radiation given by a black body depends on the body temperature and does not depend on the nature of the interior materials. Like Parachor, Rheochor is additive and constitutive. GM Baso4 cont X b) / 233 GM S. Two questions to be answered from the group 'A' and three questions to be answered
from the group 'B' Group: A atomic structure: black body radiation and quitical theory of Planck, wave particles, duality for the Equace of Elém and De-Broglie, Experimental Verification of DAVISSION and GERMER Experiment, BROGLIE wave associated with "Rbita Bohr in H-Atom, principle of Heisenberg uncertainty and his importance. Thus, the
SNCL4 becomes instable. The acid molecular mass is then calculated from the mass of the silver salt and the mass of the silver metallic residue obtained from it. The change of heat involved in the process is considered. Formation and chemical reactions of (i) unsaturated monocarboxylic acids and (II) dicarboxylic acids. Name all metals of the alkaline
earth and write down your atomic number and electrical configuration. Discuss its usefulness in the elucidation of the structure of the molems. Preparation and properties of (i) ethylene glycol (II) Glycerol and (iii) Alllyl at LCOOL. Phase Equilibrium: Phase rule and the definition of terms, involved in it, a component system. Therefore, Z = (8 x 1/8) +
(6 x 1/2) = 4 So, the number of particles per unit unit unit CCC is 4. Prove that CP CV = R For an ideal 3-Gás 3. That is, as the energy distributed between different wavelengths issued by a black body. The amount of heat needed to alter its temperature at a degree of a substance in constant volume. 8. Let's consider the three axes, OY, Oz that are cut
by a Crystal Face ABC at OA, OB and OC distance from the origin. The known weight of the organic compound is heated with steaming natric acid in the carius tube, sulfur present in the oxidized organic compound to the sulfuric acid. The first K-ionization potential is less than the CA, but the second potential of ionization is high. A 2S2 Elémera goes
to orbital 2p. 2p. + H2O -----> H3BO3 + NH3 + H2 Struture: In Borazine, both boron and nitrogen is hybridized SP2. Which of the following molems do not follow the rule of Truton A. Insolitable in Ammon Indicates iodine. What you mean by steam pressure from a liquid. Temperature, press, boiling point, fusion point, etc. are the example of intensive
properties. Q "x. ¥ Â ¥ ¥" / 2m to x. Sparently solid in Ammania Indicates Bromine Yellow PPT. 4. V â € ¥ h / 4th Mquestion: A particle is moving with a constant moment. Among the following pairs have a high moment dipolo i. Explain the deviation of the actual gas behavior. The difference between the sum of the internal energy of the products and
the sum of the internal energies of the reagents at a given temperature and constant volume is called the heat of the reaction in the constant volume at a given temperature. The vaporization entropy of most of the liquids at their boiling points is almost the same and has the value between 85-88 j / k.mol Q. of BASO4 is then washed, Dry and inflamed
to get the constant wt. If the temperature of a glass of water is 25oc, each and every drop of water in this glass has the temperature of 25oc. The temperature in which normal gases begin to behave as ideal gases (due to the absence of attractive and repulsive forces at this particular temperature). For example, ethanol combustion, Â§1366.8 kJ / mol,
is the amount of heat produced when a soft ethanol suffers a complete combustion 25 ° C and 1 pressure atmosphere . The heat of the reaction can be defined as the difference of enthalpy or heat content, "between the products and reagents. 6. What is the Boyle temperature. The short response type questions are recommended. Thus, in case of PO,
the force of the interathymical attraction of Vander Waal is less than you. Consequently, the point of fusion and boil of po is less than you. # Effect of inert pair, influence the and and sod oÉÃsiverp e oÉÃ§Ãacilpxe an oÉÃ§Ãacilpa e acid³Ãirep alebat me saicnªÃdnet saus ,ortcele edadivitagen e acin´Ârtele edadinifa ,oÉÃ§Ãazinoi ed laicnetop ,acin´Âi e
acim´Âta iidar :sacid³Ãirep sedadeirporP .otinifni on oxie mu ratroc edop ecaf adad amu uo is ertne selpmis oÉÃ§Ãroporp amu mªÃt sele erbos setnerefid sotsor ed seµÃ§Ãatpecretni sa ,sodihlocse etnemadauqeda majes soxie so es ,euq es-uocifireV .D ocir³Ãfsof odicÃ .otafsof arap odadixo missa ©Ã aicn¢Âtsbus an etneserp orofs³Ãf O .savisnetni
sedadeirporp ed sadamahc oÉÃs ametsis on etneserp aicn¢Âtsbus ad edaditnauq ad medneped oÉÃn euq sedadeirporP . arutarepmet adanimreted amu a etnatsnoc oÉÃsserp Ã oÉÃ§Ãaer ad rolac ed odamahc ©Ã etnatsnoc oÉÃsserp e arutarepmet adanimreted amu a setnegaer ed saiplatne ed amos a e sotudorp ed saiplatne sod amos a ertne
aŞÃnerefid A .5891 me oditucsid iof otsI CP ed ocolb ed sotnemele 1 :lepaP .tw[= odic¡Ã ed N701 seralom assaM)2x / 1X 801N(= n1 + n801«Â)2x / 1x 801n(=)h ed acim´Âta assam(n +)ga ed acim´Âta assam(n -)las-ga ed ralom assam(n = ralom assam ,)n(omsisab ed ocis¡Ãgilop ocilÃxobrac odic¡Ã arap 701)2x / 1x 801(= 1 + 801)2x / 1x 801(
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2x = acil¡Ãtem atarp ed g 1x = odic¡Ã essed atarp ed las ed osep = odicehnocsed ocilÃxobrac odic¡Ã ed oseP :solucl¡ÃC .A ©Ã notruT ed arger ad aivsed es euq odiuqÃl O .orgen oproc o erbos setniuges siel saud sa s´Âporp ffohcriK ,4581 mE .anilacla azerutan ed oÉÃs sasouqa seµÃ§Ãulos saus e siarenim omoc arreT oÉÃs selE .secidnI ssieW e relliM -
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Anacheidism are used to determine the structure of molecules and the nature of labels. The energy emitted or absorbed is not continuous, but is in the form of packages called quanta. Enthalpy calculation of ³ 3. Can be used for treatment of osteoporosis due to your preferred location. me. MCO3 —> Mo + CO2, then the correct order of stability of
the carbonates of group IIA A© BACO3> SRCO3> CACO 3> MGCO3> Alley3 Strong NaoH µ slowly attack, while the fused Naoh mark has made it readily berylar and releases the hydrogans. Press. 26.4 Ã¢ â¬ 67,6 k cal. How do you not determine the molecular mass of a rich carboxylic acid per mÃ ©all of silver salt. Organic compound + HNO3 â
'CO2 + H2 + H2SO4 + BACL2 + BACL2' BASO4 + 2hcl leave wt. Thus, the solubility of metal sulfates of the alkaline earth diminishes the group mainly due to the decrease in hydration energies from being + 2 to BA + 2. These properties are additive. de AG) x 108 Â"107] Acid solution basicism: 0,463 g of silver metilic acid left by silver salt = 0,759
g 216 g of silver metilic acid left by silver salt = (0,759 / 0,463) x 216 = 354 g, molecular mass of DibAbic acid = (mol. Classroom course: Xi-XII (CBSE & ICSE) Ã ", B.Sc. All the Parts (Hon's & Sub.)", M.Sc. ", Iit-Jee ", Neet", Iit-Jam ", PortÃ", Pat ", Define", Csir ", BioquMica | The warm water of microbiology in a sealed tube An example of this system.
Then, H2SO4 H is stronger than H2SO3 Borazine, also known as Borazol, There is a polar inorgÃ³nic compound with the humic formula B3N3H6. Therefore, Ã¢ â¬ z = (8 x 1/8) + (2 x 1/2) = 2 so, The number of parts per units of units for cÃ© unitÃaÃaÃ ~> O A, B and C can either n to be perpendicular. The proportion of p for pc, v to you and t for tc
and t for tc tc Press (PR), Volume (VR) and reduced temperature (TR), respectively. Inorganic chemistry role -1st 1. Therefore, the enthalpy of neutralization, the "h is always negative. Estimation of nitrogen and sulfur in organic compounds. Fulerenes are active molemats. The solid substance is heated with an oxidizing agent, such as the
concentrated natric acid or a mixture of sodium carbonate and potassium nitrate. P2O5 Write down all possible isothermages from CH3CH2CH2CH2NH2. That is why earthly alkaline metals are not obtained by electrifying their aqueous solutions. 2 = Hâ "p Â Â Â Â Â Â Â Â Â Â Â € The power of the network decreases in the group due to the increase in
the size of the section of the earthly alkali metals, while the hydration energy of the section remains almost unchanged. À to the ivors at the om to the Ã ã to the ã ã ã b2h6 + 2koh = K2 (BH2OH) 2 + H2 3. KOPP In 1842 states that The molar volume of a liquid at its boiling point is equal to the sum of the actual volume of its constituent articles.
Physical Meaning of Vander Waal B: Vander Waals 'B' constant represents the effective size of the molems. (or MS-volume excluded by a particle). After the combination, the system temperature increase is observed in the thermometer and the heat of the combustion can be calculated from the heat gained by the water and the heat. 3. B. Indicate the
hybridization, structure and form of PCl 3 and PCl5 1. If the molecular weight and the density of a substance are known, the molar volume can be easily calculated. Network energy varies as LIF> NAF> KF> RBF> CSF III. Five questions to be answered. Physical Chemistry Paper-1st 1. The law of rational investments was deducted by Hauy. This
method is based on the fact that, when an organic compound contained nitrogen is heated with Sulfuric acidnitrog³nitrog in it is quantitatively converted to sulfate. This apparatus was designed by Berthelot in 1881 to measure the heat from the combustion of organic compounds. In the increasing order of stability 1,3-pentadiene, 1,4-pentadiene,
benzene explain nucleofoile, electrophile and free radical What you don't mean by the SN1 and SN2 region I give an example of electrophalic substitution reaction join That's Maxbrain's mom for study. V ml It involves the following steps: Substance is strongly heated with NaCn Nacn Water extract is boiled with sulfate iron solution. The mol ©cula
C60 has two lengths of labels - the ring µ6 links: 6 can be considered as "double bonds" and are shorter than µ bonds 6: 5. Fulerene A© the third form of carbon material ³ graphite and diamond whose molÃ ©cula consists of 20 anÃ ©is hexagonal and 12 pentagonal as the basis of a cage structure of icosohedral symmetry. This is explained by the
polarizing ability of the non-believers of the IIA group. 7. Preparation, Properties and Estimation of Urea. Aldeans and ketones: nomenclature, carbonyl group structure, m ©all general preparation, properties of aldos and ketones, an introduction to aldos and unsaturated ketones of Â± ". SN & PB shows +2 and +4 oxidation states while TL shows +3
and +1 oxidation state due to inert pair effect. These properties are non-additive. We know that h = e + pv or, h = e + rt (as pv = rt for a mole) differentiating the equation above wrt, n³s we receive - dh / dt = de / dt + r (dt / dt) or, CP = CV + R or, CP Â"CV = R TambÃ© m A© called the energy conservation principle. What's the transport number?
Please justify. Orgµnic Reaction Types, Energy Consideration with Reference activation energy and transition state. Why is acetylene more acid than ethylene? Limiting Radius Ratio b. Then, the inert pair effect becomes more .'et' .'et' euq od 'op' arap grebnesieH .stnatsnoc slaaw red nav fo smret ni noitaulave rieht dna stnatsnoc lacitirc
,nonemonehp lacitirc ,etats fo noitauque laaW red naV fo noitavireD :etatS suoesaG 'B' puorg morf derewsna eb ot snoitseuq eerht dna 'A' puorg morf derewsna eb ot snoitseuq owT .elur s F ,snoi fo ytilibiziralop dna rewop gniziralop ,dilos cinoi fo ytilibulos ygrene noitavlos ,elcyc rebaH nroB ,ygrene ecittaL :sdiloS cinoI .secidnI lanoitaror fo waL eht
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ralucelomretni fo edutingam eht stneserper 'a' tnatsnoc laaW rednaV :a stnatsnoC laaW rednaV fo ecnacifingiS lacisyhP Â Â³Ãâlom retil :'b' fo tinU Â2ÃλÂ"Âλâm Â2Ãλtiltil mta :'a' fo tinU emulov ralom = V ,tnatsnoc laaW rednaV = b & a ,erutarepmet etulosba = T ,tnatsnoc sag lasrevinu = R ,erusserp = P :erehW .3 R = VC ÂTt evorP .puorg eht nwod
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ÃτÀLÂ ÃτlÂ ÃτTEÂ ÃτEÂ ÃτTE3Â + SbP ÂTE³EE33 + S2N + 2)OOC3HC(bPÂ ÃτÂ Â ÃτÂ Â ÃτÂ Â ÃτÂ ÃTE313333334333333333333333444444433334433 in .ametsis .ametsis etsed olpmexe mu ©Ã opoc mun etneuq aug¡Ã A .d oipÃcnirP

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